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1. What is the vapor pressure of water when it is boiling at 30.0°C?
A) .2 kPa B) 5 kPa
C) 45 kPa D) 101.3 kPa
 2. What is the vapor pressure of water at 105°C?
A) .60 kPa B) 101.3 kPa
C) 120 kPa D) 145 kPa
 3. When the vapor pressure of water is 30 kPa, the temperature of the water is
A) 20°C B) 40°C
C) 70°C D) 100°C
 4. Which sample of water has the greatest vapor pressure?
A) 100 ml at 20°C B) 200 ml at 25°C
C) 20 ml at 30°C D) 40 ml at 35°C
 5. According to Reference Table *H*, what is the vapor pressure of propanone at 45°C?
A) 22 kPa B) 33 kPa
C) 70 kPa D) 98 kPa
 6. When the temperature of a sample of water is changed from 45°C to 70.°C, the change in its vapor pressure is
A) 1.0 kPa B) 20. kPa
C) 25 kPa D) 101.3 kPa
 7. At 298 K, the vapor pressure of H₂O is less than the vapor pressure of CS₂. The best explanation for this is that H₂O has
A) larger molecules
B) a larger molecular mass
C) stronger ionic bonds
D) stronger intermolecular forces
 8. Based on Reference Table *H*, which sample has the highest vapor pressure?
A) water at 20°C B) water at 80°C
C) ethanol at 50°C D) ethanol at 65°C
 9. The vapor pressure of a liquid is 0.92 atm at 60°C. The normal boiling point of the liquid could be
A) 35°C B) 45°C C) 55°C D) 65°C
 10. Using your knowledge of chemistry and the information in Reference Table *H*, which statement concerning propanone and water at 50°C is true?
A) Propanone has a higher vapor pressure and stronger intermolecular forces than water.
B) Propanone has a higher vapor pressure and weaker intermolecular forces than water.
C) Propanone has a lower vapor pressure and stronger intermolecular forces than water.
D) Propanone has a lower vapor pressure and weaker intermolecular forces than water.
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